

Formation of carbonic acid equation

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P305+P351+P338 NFPA 704 (fire diamond) 1 3 0 Flash point 22 °C (72 °F; 295 K) Autoignitiontemperature 371 °C (700 °F; 644 K) Explosive point 2.2–13.7%^[2] Lethal dose or concentration (LD, LC): LD50 (median dose) 2800 mg/kg (rabbit, oral)[4] 1870 mg/kg (rat, oral)[5] NIOSH (US health exposure limits): PEL (Permissible) TWA 200 ppm (500 mg/m3)[2] REL (Recommended) TWA 200 ppm (500 mg/m3) ST 250 ppm (625 mg/m3) [skin][2] IDLH (Immediate danger) 800 ppm[2] Related compounds Related compounds Propionalsopropyl alcoholPropanamineEthanolButanol Supplementary data page 1-Propanol (data page) Except where otherwise noted, data are given for materials in their standard state (at 25 °C [77 °F], 100 kPa). N verify (what is YN ?) Infobox references Chemical compound Propan-1-ol (also propanol, n-propyl alcohol) is a primary alcohol with the formula CH3CH2CH2OH and sometimes represented as PrOH or n-PrOH. It is a colorless liquid and an isomer of 2-propanol. It is formed naturally in small amounts during many fermentation processes and used as a solvent in the pharmaceutical industry, mainly for resins and cellulose esters, and, sometimes, as a disinfecting agent. Chemical properties Some example reactions of propan-1-olPropan-1-ol shows the normal reactions of a primary alcohol. Thus it can be converted to alkyl halides; for example red phosphorus and iodine produce n-propyl iodide in 80% yield, while PCl3 with catalytic ZnCl2 gives n-propyl chloride. Reaction with acetic acid in the presence of an H2SO4 catalyst under Fischer esterification conditions gives propyl acetate, while refluxing propanol overnight with formic acid alone can produce propyl formate in 65% yield. Oxidation of propan-1-ol with Na2Cr2O7 and H2SO4 gives a 36% yield of propionaldehyde, and therefore for this type of reaction higher yielding methods using PCC or the Swern oxidation are recommended. Oxidation with chromic acid yields propionic acid. Preparation Propan-1-ol is manufactured by catalytic hydrogenation of propionaldehyde. Propionaldehyde is produced via the oxo process by hydroformylation of ethylene using carbon monoxide and hydrogen in the presence of a catalyst such as cobalt octacarbonyl or a rhodium complex.[6] H2C=CH2 + CO + H2 → CH3CH2CH=O CH3CH2CH=O + H2 → CH3CH2CH2OH A traditional laboratory preparation of propan-1-ol involves treating n-propyl iodide with moist Ag2O. Safety Propan-1-ol is thought to be similar to ethanol in its effects on the human body, but 2–4 times more potent. Oral LD50 in rats is 1870 mg/kg (compared to 7060 mg/kg for ethanol). It is metabolized into propionic acid. Effects include alcoholic intoxication and high anion gap metabolic acidosis. As of 2011, one case of lethal propan-1-ol poisoning was reported.[7] Propan-1-ol as fuel Propan-1-ol has high octane number and is suitable for engine fuel usage. However, propanol is too expensive to use as a motor fuel. The research octane number (RON) of propanol is 118, and anti-knock index (AKI) is 108.[8] References ^ Favre HA, Powell WH (2014). Nomenclature of Organic Chemistry: IUPAC Recommendations and Preferred Names 2013 (Blue Book). Cambridge: The Royal Society of Chemistry. p. 61. doi:10.1039/9781849733069. ISBN 9780854041824. ^ a b c d e NIOSH Pocket Guide to Chemical Hazards. "#0533". National Institute for Occupational Safety and Health (NIOSH). ^ Pal A, Gaba R (2008). "Volumetric, acoustic, and viscometric studies of molecular interactions in binary mixtures of dipropylene glycol dimethyl ether with 1-alkanols at 298.15 K". *J. Chem. Thermodyn.* 40 (5): 818–828. doi:10.1016/j.jct.2008.01.008. ^ Propan-1-ol toxicity ^ "n-Propyl alcohol". Immediately Dangerous to Life or Health Concentrations (IDLH). National Institute for Occupational Safety and Health (NIOSH). ^ Papa AJ (2011). "Propanols". Ullmann's Encyclopedia of Industrial Chemistry. 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External links International Chemical Safety Card 0553 NIOSH Pocket Guide to Chemical Hazards Retrieved from " is a process used by plants and other organisms to convert light energy into chemical energy that, through cellular respiration, can later be released to fuel the organism's activities.Some of this chemical energy is stored in carbohydrate molecules, such as sugars and starches, which are synthesized from carbon dioxide and water - hence the name ... 1) CO 2 + H 2 O – Carbonic anhydrase H 2 CO 3

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