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## Xeo3 lewis structure

XEO3 - Trioxido Xenon: First Draw The Lewis Dot Structure: Electronic Geometry: Tetrahedral. Hybridization: SP3 then draws the 3D molecular structure using VSEPR rules: Decision: XEOO3 molecular geometry is trigonal pyramidal with asymmetrical charge distribution on the central atom. Therefore this molecule is polar. Xenon trioxide on Wikipedia. Back to molecular geometries and polarity tutorial: molecular geometry and polarity tutorials. For homework Help in mathematics, chemistry and physics: [www.tutor-homework.com](http://www.tutor-homework.com). Postby Kevin Wright Á, Á »SAT 24 October 2015 9:45 PM For the Lewis structure of Xenon trioxide, it would be drawn with 3 double ties and a solitary pair of electron on the xenon. Yes, the formal charge is the decisive factor for this. If they were all the individual ties, it would satisfy the octet rule, but would result in a formal  $3+ \text{ } \ddot{\text{e}}$







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